PPSC Lecturer Chemistry | BS-17 — GUESS PAPER 2026 (Set 2 of 2) · Predicted Questions

Body: Punjab Public Service Commission (PPSC) · **Post:** Lecturer Chemistry | BS-17 · **Set:** 2 of 2 · **Questions:** 100 predicted — 0 repeated questions, 32 verified facts predicted to recur, 68 syllabus anchors · **Price:** Free

GUESS PAPER — predicted questions for the upcoming exam, NOT a past paper. Every question below is a prediction: it is here because the same question, the same fact, or the same topic appeared in verified PPSC Lecturer Chemistry | BS-17 past papers. **No pass guarantee.** A guess paper raises the odds; it does not promise a result. Answers are the best-supported answers, not exam keys.

[C]-HEAVY PAPER — thin verified pool. 68 of the 100 questions below are syllabus anchors on this post's recurring topics, because the verified archive for this post holds only 66 items. Those questions are flagged [C] on every one of them: they are predictions about the syllabus, not verified facts. Study the topic, not just the answer.

> **How to use this paper:** study the basis under each question, not just the answer — the real paper will ask the same topics in different words. Revise the underlying topic, not only the four options. Do not quote any item as "asked in the 2026 exam" — nothing here has been asked yet.

Question 1 Chemistry

A complex with coordination number 4 shows no cis-trans isomerism. What must its structure be?

- A. Linear
- B. Trigonal planar

- tahedral ✓ **Predicted answer**
- D. None of these

Why this answer? A four-coordinate complex with no cis-trans isomerism must be tetrahedral.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q38 (verified item — correct answer: Tetrahedral).

Question 2 Chemistry

What is the coordination number of Pt in PtCl(NO2)(NH3)4?

- A. 5
- B. 4
- C. 6 ✓ **Predicted answer**
- D. None of these

Why this answer? Pt is bonded to Cl, NO2 and four NH3 ligands: 1 + 1 + 4 = 6.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q39 (verified item — correct answer: 6).

Question 3 Chemistry

Solvent extraction is controlled by which law?

- A. Nernst distribution law ✓ **Predicted answer**
- B. Henry's law
- C. Kirchhoff's law
- D. None of these

Why this answer? Solvent extraction is controlled by the Nernst distribution law.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q40 (verified item — correct answer: Nernst distribution law).

4 Chemistry

Compared with a single extraction, repeated extractions are more:

- A. Reliable
- B. Better
- C. None of these
- D. Efficient ✓ **Predicted answer**

Why this answer? Repeated extractions are more efficient than a single extraction.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q41 (verified item — correct answer: Efficient).

Question 5 Chemistry

What pH is maintained when calcium is titrated with EDTA?

- A. None of these (pH 10) ✓ **Predicted answer**
- B. pH 8
- C. pH 6
- D. pH 7

Why this answer? Calcium is titrated with EDTA at pH 10, which is not among the listed values.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q43 (verified item — correct answer: None of these).

Question 6 Chemistry

In the titration where MnO4- is reduced to Mn2+, the endpoint is shown by:

- A. Disappearing pink colour
- B. A redox indicator
- C. None of these

ppearing pink colour ✓ **Predicted answer**

Why this answer? Permanganate is its own indicator: the endpoint is shown by the appearing pink colour.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q46 (verified item — correct answer: Appearing pink color).

Question 7 Chemistry

How is the endpoint of a redox titration determined?

- A. Phenolphthalein
- B. Eriochrome Black T
- C. None of these ✓ **Predicted answer**
- D. Methyl orange

Why this answer? None of the listed indicators is a redox indicator; the endpoint is found with a redox indicator or a self-indicator.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q47 (verified item — correct answer: None of these).

Question 8 Chemistry

At about which wavenumber does a carbonyl group absorb in the infrared?

- A. 1700 cm-1 ✓ **Predicted answer**
- B. 1600 cm-1
- C. 2950 cm-1
- D. None of these

Why this answer? A carbonyl group absorbs at about 1700 cm-1.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q49 (verified item — correct answer: 1700).

9 Chemistry

Monochromators are used in spectroscopy for:

- A. Separation of mixtures
- B. Resolving wavelengths ✓ **Predicted answer**
- C. Solving equations
- D. None of these

Why this answer? Monochromators resolve (separate) wavelengths.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q51 (verified item — correct answer: Resolving).

Question 10 Chemistry

Which mode of energy transfer needs no physical medium?

- A. Convection
- B. Radiations ✓ **Predicted answer**
- C. Conduction
- D. None of these

Why this answer? Radiation transfers energy without any physical medium.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q52 (verified item — correct answer: Radiations).

Question 11 Chemistry

Which is the descending order of gases for global warming effect?

- A. CO2, CH4, H2O, ClFCs ✓ **Predicted answer**
- B. CO2, ClFCs, H2O, CH4
- C. CO2, H2O, CH3, ClFCs
- D. None of these

swer? The descending order is CO2, CH4, H2O, ClFCs.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q53 (verified item — correct answer: CO2, CH4, H2O, ClFCs).

Question 12 Chemistry

Mass transport in the environment is mainly the work of which parts?

- A. Lithosphere and hydrosphere
- B. Lithosphere and atmosphere
- C. None of these
- D. Atmosphere and hydrosphere ✓ **Predicted answer**

Why this answer? The atmosphere and hydrosphere are mainly responsible for mass transport.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q54 (verified item — correct answer: Atmosphere/hydrosphere).

Question 13 Chemistry

Acid rain occurs mainly because of which gases?

- A. NO2 and SO2 ✓ **Predicted answer**
- B. CO2 and NO2
- C. SO2 and CO2
- D. None of these

Why this answer? Acid rain is caused mainly by NO2 and SO2.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q55 (verified item — correct answer: NO2/SO2).

14 Chemistry

Who discovered the formation of ozone?

- A. Dubsen
- B. Jonathan
- C. None of these
- D. Schonbein ✓ **Predicted answer**

Why this answer? Schonbein discovered ozone.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q56 (verified item — correct answer: Schonbein).

Question 15 Chemistry

Sulphuric acid as a byproduct is obtained from which process?

- A. Elemental sulfur mining
- B. Oxidation reduction
- C. None of these
- D. Pyrolysis of pyrites ✓ **Predicted answer**

Why this answer? Sulphuric acid is a byproduct of the pyrolysis of pyrites.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q57 (verified item — correct answer: Pyrolysis of pyrites).

Question 16 Chemistry

Which of the following statements about mercury, lead and cadmium is true?

- A. All three are red-list hazardous metals ✓ **Predicted answer**
- B. Only mercury is hazardous
- C. Only lead is hazardous
- D. None of them is hazardous

swer? Mercury, lead and cadmium are all red-list hazardous metals.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q59 (verified item — correct answer: All of the above).

Question 17 Chemistry

Baeyer's reagent is decolourised and a vicinal diol forms. What functional group is present?

- A. Alkane
- B. Cyclohexane
- C. None of these
- D. Alkene ✓ **Predicted answer**

Why this answer? Baeyer's reagent is decolourised by alkenes, forming a vicinal diol.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q61 (verified item — correct answer: Alkene).

Question 18 Chemistry

2,3-dibromobutane (CH3CHBrCHBrCH3) with zinc in ethanol gives which product?

- A. Alkane
- B. Alcohol
- C. Alkene ✓ **Predicted answer**
- D. None of these

Why this answer? Debromination of 2,3-dibromobutane with zinc gives an alkene.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q62 (verified item — correct answer: Alkene).

19 Chemistry

Toluene is converted to benzaldehyde with chromyl chloride. What is this reaction called?

- A. Etard reaction ✓ **Predicted answer**
- B. Cannizzaro reaction
- C. Wurtz reaction
- D. None of these

Why this answer? The conversion of toluene to benzaldehyde with chromyl chloride is the Etard reaction.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q64 (verified item — correct answer: Etard reaction).

Question 20 Chemistry

On heating, beta-hydroxybutyric acid gives which product?

- A. Lactone
- B. Butyric acid
- C. None of these
- D. Crotonic acid ✓ **Predicted answer**

Why this answer? Heating beta-hydroxybutyric acid gives crotonic acid.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q66 (verified item — correct answer: Crotonic acid).

Question 21 Chemistry

How many asymmetric carbons are there in 2,3-dibromopropanol?

- A. 1 ✓ **Predicted answer**
- B. 2

- D. None of these

Why this answer? 2,3-dibromopropanol has one asymmetric carbon.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q67 (verified item — correct answer: 1).

Question 22 Chemistry

What is the name of Vitamin B2?

- A. Cobalamin
- B. Niacin
- C. None of these (Riboflavin) ✓ **Predicted answer**
- D. Thiamine

Why this answer? Vitamin B2 is riboflavin, which is not listed; cobalamin is B12, niacin is B3 and thiamine is B1.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q68 (verified item — correct answer: None of these).

Question 23 Chemistry

Iron is present in:

- A. Chlorophyll
- B. Pyridoxine
- C. Haemoglobin ✓ **Predicted answer**
- D. None of these

Why this answer? Iron is present in haemoglobin.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q70 (verified item — correct answer: Hemoglobin).

24 Chemistry

A nucleoside is composed of:

- A. Phosphate and base
- B. Phosphate and sugar
- C. None of these
- D. Base and sugar ✓ **Predicted answer**

Why this answer? A nucleoside is composed of a base and a sugar.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q71 (verified item — correct answer: Base-sugar).

Question 25 Chemistry

Which base is present in RNA but not in DNA?

- A. None of these (Uracil) ✓ **Predicted answer**
- B. Adenine
- C. Guanine
- D. Cytosine

Why this answer? Uracil is present in RNA but not in DNA; it is not listed among the options.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q72 (verified item — correct answer: None of these (Uracil)).

Question 26 Chemistry

Isomers that differ in configuration at only one carbon are called:

- A. Epimers ✓ **Predicted answer**
- B. Anomers
- C. Enantiomers

Why this answer? Isomers differing in configuration at only one carbon are epimers.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q74 (verified item — correct answer: Epimers).

Question 27 Chemistry

Which alkali is used for cold-process soap making?

- A. KOH
- B. Na2CO3
- C. None of these
- D. NaOH ✓ **Predicted answer**

Why this answer? Cold-process saponification uses NaOH.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q75 (verified item — correct answer: NaOH).

Question 28 Chemistry

How many peptide bonds are there in a decapeptide?

- A. 6
- B. 10
- C. 9 ✓ **Predicted answer**
- D. None of these

Why this answer? A decapeptide has 10 amino acids joined by 9 peptide bonds.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q76 (verified item — correct answer: 9).

29 Chemistry

Body fluids contain calcium, sodium, nitrates and sulphate as what?

- A. Electrolytes ✓ **Predicted answer**
- B. Minerals
- C. Micro minerals
- D. None of these

Why this answer? Calcium, sodium, nitrates and sulphate in the body are electrolytes.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q78 (verified item — correct answer: Electrolytes).

Question 30 Chemistry

What is the most abundant element in the human body?

- A. None of these (Oxygen) ✓ **Predicted answer**
- B. Calcium
- C. Potassium
- D. Phosphorus

Why this answer? Oxygen is the most abundant element in the human body; it is not listed.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q79 (verified item — correct answer: None of these).

Question 31 Chemistry

Metabolic reactions in the body are catalysed by:

- A. Hormones
- B. Vitamins
- C. None of these

ymes ✓ **Predicted answer**

Why this answer? Metabolic reactions are catalysed by enzymes.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q80 (verified item — correct answer: Enzymes).

Question 32 Chemistry

Light of which wavelengths is able to drive a photochemical reaction?

- A. 2000 to 12000 angstrom ✓ **Predicted answer**
- B. 500 to 1000 angstrom
- C. 1000 to 1500 angstrom
- D. None of these

Why this answer? Photochemical reactions are brought about by radiation of about 2000 to 12000 angstrom.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q33 (verified item — correct answer: 2000 to 12000 A).

Question 33 Chemistry

How many moles are there in 9 g of a substance of molar mass 3 g/mol?

- A. 4
- B. 2
- C. 3 ✓ **Predicted answer**
- D. 6

Why this answer? Moles = mass / molar mass = 9 / 3 = 3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

34 Chemistry

How many moles are there in 15 g of a substance of molar mass 5 g/mol?

- A. 4
- B. 3 ✓ **Predicted answer**
- C. 2
- D. 6

Why this answer? Moles = mass / molar mass = 15 / 5 = 3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

Question 35 Chemistry

How many moles are there in 24 g of a substance of molar mass 6 g/mol?

- A. 5
- B. 3
- C. 8
- D. 4 ✓ **Predicted answer**

Why this answer? Moles = mass / molar mass = 24 / 6 = 4.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

Question 36 Chemistry

How many moles are there in 30 g of a substance of molar mass 5 g/mol?

- A. 7
- B. 5
- C. 6 ✓ **Predicted answer**
- D. 12

swer? Moles = mass / molar mass = 30 / 5 = 6.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

Question 37 Chemistry

What is the molarity of a solution containing 4 moles of solute in 2 litres?

- A. 3 M
- B. 2 M ✓ **Predicted answer**
- C. 1 M
- D. 4 M

Why this answer? Molarity = moles / volume = 4 / 2 = 2 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 38 Chemistry

What is the molarity of a solution containing 10 moles of solute in 2 litres?

- A. 6 M
- B. 4 M
- C. 7 M
- D. 5 M ✓ **Predicted answer**

Why this answer? Molarity = moles / volume = 10 / 2 = 5 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 39 Chemistry

What is the molarity of a solution containing 9 moles of solute in 3 litres?

- B. 2 M
- C. 5 M
- D. 3 M ✓ **Predicted answer**

Why this answer? Molarity = moles / volume = 9 / 3 = 3 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 40 Chemistry

What is the molarity of a solution containing 12 moles of solute in 4 litres?

- A. 3 M ✓ **Predicted answer**
- B. 4 M
- C. 2 M
- D. 5 M

Why this answer? Molarity = moles / volume = 12 / 4 = 3 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 41 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-5 M?

- A. 6
- B. 4
- C. 7
- D. 5 ✓ **Predicted answer**

Why this answer? pH = -log[H+] = -log(10^-5) = 5.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs

urer Chemistry sitting.

Question 42 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-6 M?

- A. 7
- B. 5
- C. 6 ✓ **Predicted answer**
- D. 8

Why this answer? pH = -log[H+] = -log(10^-6) = 6.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 43 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-7 M?

- A. 8
- B. 6
- C. 7 ✓ **Predicted answer**
- D. 9

Why this answer? pH = -log[H+] = -log(10^-7) = 7.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 44 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-8 M?

- A. 9
- B. 8 ✓ **Predicted answer**

- D. 10

Why this answer? pH = -log[H+] = -log(10^-8) = 8.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 45 Chemistry

A gas at 5 atm occupies 10 L. At constant temperature its volume becomes 2 L. What is the new pressure?

- A. 26 atm
- B. 24 atm
- C. 25 atm ✓ **Predicted answer**
- D. 27 atm

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 5 x 10 / 2 = 25 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 46 Chemistry

A gas at 8 atm occupies 4 L. At constant temperature its volume becomes 8 L. What is the new pressure?

- A. 5 atm
- B. 3 atm
- C. 4 atm ✓ **Predicted answer**
- D. 6 atm

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 8 x 4 / 8 = 4 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in

er Chemistry sitting.

Question 47 Chemistry

A gas at 2 atm occupies 6 L. At constant temperature its volume becomes 12 L. What is the new pressure?

- A. 2 atm
- B. 0 atm
- C. 3 atm
- D. 1 atm ✓ **Predicted answer**

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 2 x 6 / 12 = 1 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 48 Chemistry

A gas at 10 atm occupies 5 L. At constant temperature its volume becomes 5 L. What is the new pressure?

- A. 11 atm
- B. 9 atm
- C. 12 atm
- D. 10 atm ✓ **Predicted answer**

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 10 x 5 / 5 = 10 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 49 Chemistry

What is the molar mass of HCl?

- B. 36.5 g/mol ✓ **Predicted answer**
- C. 37 g/mol
- D. 40 g/mol

Why this answer? The molar mass of HCl is 36.5 g/mol.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molar mass recurs in every Lecturer Chemistry sitting.

Question 50 Chemistry

What is the molar mass of NH3?

- A. 16 g/mol
- B. 18 g/mol
- C. 17 g/mol ✓ **Predicted answer**
- D. 14 g/mol

Why this answer? The molar mass of NH3 is 17 g/mol.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molar mass recurs in every Lecturer Chemistry sitting.

Question 51 Chemistry

What is the molar mass of CO2?

- A. 42 g/mol
- B. 46 g/mol
- C. 44 g/mol ✓ **Predicted answer**
- D. 28 g/mol

Why this answer? The molar mass of CO2 is 44 g/mol.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molar mass recurs

urer Chemistry sitting.

Question 52 Chemistry

Which element has the chemical symbol K?

- A. Potassium ✓ **Predicted answer**
- B. Krypton
- C. Phosphorus
- D. Platinum

Why this answer? K is the symbol for potassium.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

Question 53 Chemistry

Which element has the chemical symbol Ag?

- A. Argon
- B. Silver ✓ **Predicted answer**
- C. Aluminium
- D. Gold

Why this answer? Ag (from argentum) is the symbol for silver.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

Question 54 Chemistry

Which element is the most abundant in the Earth's crust?

- A. Silicon
- B. Oxygen ✓ **Predicted answer**

- D. Iron

Why this answer? Oxygen is the most abundant element in the Earth's crust.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic table recurs in every Lecturer Chemistry sitting.

Question 55 Chemistry

Which gas makes up most of the air?

- A. Nitrogen ✓ **Predicted answer**
- B. Oxygen
- C. Argon
- D. Carbon dioxide

Why this answer? Nitrogen makes up about 78 percent of the air.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gases recurs in every Lecturer Chemistry sitting.

Question 56 Chemistry

Which acid is present in citrus fruits?

- A. Acetic acid
- B. Citric acid ✓ **Predicted answer**
- C. Tartaric acid
- D. Lactic acid

Why this answer? Citrus fruits contain citric acid.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

57 Chemistry

Which base is used in soap making?

- A. Sodium hydroxide ✓ **Predicted answer**
- B. Calcium hydroxide
- C. Ammonium hydroxide
- D. Potassium nitrate

Why this answer? Sodium hydroxide is used in soap making.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

Question 58 Chemistry

What is the pH of lemon juice?

- A. About 7
- B. About 10
- C. About 12
- D. About 2 ✓ **Predicted answer**

Why this answer? Lemon juice is acidic, with pH about 2.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 59 Chemistry

Which of the following is a weak acid?

- A. Hydrochloric acid
- B. Acetic acid ✓ **Predicted answer**
- C. Sulphuric acid
- D. Nitric acid

swer? Acetic acid is a weak acid.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

Question 60 Chemistry

Which gas is released when a carbonate reacts with an acid?

- A. Hydrogen
- B. Carbon dioxide ✓ **Predicted answer**
- C. Oxygen
- D. Nitrogen

Why this answer? Carbonates react with acids to release carbon dioxide.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

Question 61 Chemistry

Which metal is used for making aeroplane bodies because it is light and strong?

- A. Iron
- B. Copper
- C. Aluminium ✓ **Predicted answer**
- D. Zinc

Why this answer? Aluminium is light and strong and is used in aeroplanes.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: metals recurs in every Lecturer Chemistry sitting.

Question 62 Chemistry

Which metal is used to galvanise iron?

: ✓ **Predicted answer**

- B. Tin
- C. Copper
- D. Nickel

Why this answer? Zinc is used to galvanise iron.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: metals recurs in every Lecturer Chemistry sitting.

Question 63 Chemistry

Which allotrope of carbon is the hardest natural substance?

- A. Diamond ✓ **Predicted answer**
- B. Graphite
- C. Charcoal
- D. Fullerene

Why this answer? Diamond is the hardest natural substance.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: non-metals recurs in every Lecturer Chemistry sitting.

Question 64 Chemistry

What is the chemical formula of water?

- A. H2O ✓ **Predicted answer**
- B. CO2
- C. H2O2
- D. O2

Why this answer? Water is H2O.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common

recurs in every Lecturer Chemistry sitting.

Question 65 Chemistry

What is the chemical formula of washing soda?

- A. NaHCO3
- B. NaOH
- C. CaCO3
- D. Na2CO3 ✓ **Predicted answer**

Why this answer? Washing soda is sodium carbonate, Na2CO3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 66 Chemistry

What is the chemical formula of limestone?

- A. CaO
- B. CaCO3 ✓ **Predicted answer**
- C. Ca(OH)2
- D. CaCl2

Why this answer? Limestone is calcium carbonate, CaCO3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 67 Chemistry

Which gas is used to extinguish fire?

- A. Carbon dioxide ✓ **Predicted answer**
- B. Oxygen

rogen

- D. Nitrogen

Why this answer? Carbon dioxide is used in fire extinguishers.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gases recurs in every Lecturer Chemistry sitting.

Question 68 Chemistry

Which element has the lowest electronegativity?

- A. Fluorine
- B. Caesium ✓ **Predicted answer**
- C. Lithium
- D. Sodium

Why this answer? Caesium has the lowest electronegativity.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic trends recurs in every Lecturer Chemistry sitting.

Question 69 Chemistry

Which group of the periodic table contains the halogens?

- A. Group 18
- B. Group 17 ✓ **Predicted answer**
- C. Group 16
- D. Group 1

Why this answer? The halogens are in group 17.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic table recurs in every Lecturer Chemistry sitting.

70 Chemistry

Which of the following is a halogen?

- A. Argon
- B. Chlorine ✓ **Predicted answer**
- C. Oxygen
- D. Sulphur

Why this answer? Chlorine is a halogen.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic table recurs in every Lecturer Chemistry sitting.

Question 71 Chemistry

What is the atomic number of oxygen?

- A. 6
- B. 16
- C. 8 ✓ **Predicted answer**
- D. 10

Why this answer? Oxygen has atomic number 8.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 72 Chemistry

How many protons does a carbon atom have?

- A. 12
- B. 6 ✓ **Predicted answer**
- C. 4
- D. 8

swer? Carbon has atomic number 6, so it has 6 protons.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 73 Chemistry

Which particle has a positive charge?

- A. Electron
- B. Neutron
- C. Proton ✓ **Predicted answer**
- D. Photon

Why this answer? The proton carries a positive charge.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 74 Chemistry

Who discovered the electron?

- A. Rutherford
- B. Bohr
- C. J. J. Thomson ✓ **Predicted answer**
- D. Chadwick

Why this answer? J. J. Thomson discovered the electron.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 75 Chemistry

What is the chemical symbol for zinc?

- B. Zn ✓ **Predicted answer**
- C. Z
- D. Sn

Why this answer? Zn is the symbol for zinc.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

Question 76 Chemistry

What is the chemical symbol for gold?

- A. Ag
- B. Gd
- C. Au ✓ **Predicted answer**
- D. Go

Why this answer? Au (from aurum) is the symbol for gold.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

Question 77 Chemistry

Which element is used in pencil lead?

- A. Diamond
- B. Coal
- C. Silicon
- D. Graphite ✓ **Predicted answer**

Why this answer? Pencil lead is graphite.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: non-metals recurs

urer Chemistry sitting.

Question 78 Chemistry

What is the chemical formula of nitric acid?

- A. H2SO4
- B. HNO3 ✓ **Predicted answer**
- C. HCl
- D. HNO2

Why this answer? Nitric acid is HNO3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 79 Chemistry

Which process is used to separate salt from water?

- A. Filtration
- B. Evaporation ✓ **Predicted answer**
- C. Sublimation
- D. Distillation

Why this answer? Salt is separated from water by evaporation.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: separation techniques recurs in every Lecturer Chemistry sitting.

Question 80 Chemistry

Which acid is used in car batteries?

- A. Sulphuric acid ✓ **Predicted answer**
- B. Nitric acid

rochloric acid

- D. Acetic acid

Why this answer? Car batteries use sulphuric acid.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 81 Everyday Science

The freezing point of water is:

- A. 10 degrees Celsius
- B. 0 degrees Celsius ✓ **Predicted answer**
- C. 5 degrees Celsius
- D. minus 10 degrees Celsius

Why this answer? Water freezes at 0 degrees Celsius.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: basic physics facts recur in every Lecturer sitting.

Question 82 Computer

What does CPU stand for?

- A. Central Program Unit
- B. Computer Processing Unit
- C. Central Processor Utility
- D. Central Processing Unit ✓ **Predicted answer**

Why this answer? CPU stands for Central Processing Unit.

Why predicted [C] — Syllabus anchor: syllabus anchor — Computer: basic computer terms recur in every Lecturer sitting.

83 Computer

Which of these is an output device?

- A. Keyboard
- B. Mouse
- C. Monitor ✓ **Predicted answer**
- D. Scanner

Why this answer? A monitor is an output device.

Why predicted [C] — Syllabus anchor: syllabus anchor — Computer: basic computer terms recur in every Lecturer sitting.

Question 84 Computer

How many bits are there in a byte?

- A. 8 ✓ **Predicted answer**
- B. 4
- C. 16
- D. 32

Why this answer? There are 8 bits in a byte.

Why predicted [C] — Syllabus anchor: syllabus anchor — Computer: basic computer terms recur in every Lecturer sitting.

Question 85 Computer

What does WWW stand for?

- A. World Web Wide
- B. Wide World Web
- C. Web World Wide
- D. World Wide Web ✓ **Predicted answer**

swer? WWW stands for World Wide Web.

Why predicted [C] — Syllabus anchor: syllabus anchor — Computer: basic computer terms recur in every Lecturer sitting.

Question 86 Computer

Which key is used to delete text to the left of the cursor?

- A. Enter
- B. Backspace ✓ **Predicted answer**
- C. Tab
- D. Shift

Why this answer? Backspace deletes text to the left of the cursor.

Why predicted [C] — Syllabus anchor: syllabus anchor — Computer: basic computer terms recur in every Lecturer sitting.

Question 87 General Knowledge

What is the capital of Australia?

- A. Sydney
- B. Canberra ✓ **Predicted answer**
- C. Melbourne
- D. Perth

Why this answer? Canberra is the capital of Australia; Sydney is its largest city.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world capitals recur in every Lecturer sitting.

Question 88 General Knowledge

Which is the largest desert in the world?

ara ✓ **Predicted answer**

- B. Gobi
- C. Kalahari
- D. Thar

Why this answer? The Sahara is the largest hot desert in the world.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 89 General Knowledge

Which river is the longest in the world?

- A. Amazon
- B. Yangtze
- C. Nile ✓ **Predicted answer**
- D. Indus

Why this answer? The Nile (about 6,650 km) is generally regarded as the world's longest river.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 90 General Knowledge

Which country is home to the kangaroo?

- A. New Zealand
- B. South Africa
- C. Australia ✓ **Predicted answer**
- D. Brazil

Why this answer? The kangaroo is native to Australia.

ed [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 91 General Knowledge

What is the capital of Canada?

- A. Toronto
- B. Vancouver
- C. Ottawa ✓ **Predicted answer**
- D. Montreal

Why this answer? Ottawa is the capital of Canada.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world capitals recur in every Lecturer sitting.

Question 92 General Knowledge

Which is the smallest continent?

- A. Europe
- B. Australia ✓ **Predicted answer**
- C. Antarctica
- D. Africa

Why this answer? Australia is the smallest continent.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 93 General Knowledge

Which ocean is the largest?

- A. Atlantic

- C. Arctic
- D. Pacific ✓ **Predicted answer**

Why this answer? The Pacific is the largest ocean.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 94 General Knowledge

Which is the largest island in the world?

- A. Madagascar
- B. Borneo
- C. Sumatra
- D. Greenland ✓ **Predicted answer**

Why this answer? Greenland is the largest island in the world.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 95 General Knowledge

What is the capital of Japan?

- A. Kyoto
- B. Osaka
- C. Nagoya
- D. Tokyo ✓ **Predicted answer**

Why this answer? Tokyo is the capital of Japan.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world capitals recur in every Lecturer sitting.

96 General Knowledge

Which country has the largest population?

- A. China
- B. USA
- C. India ✓ **Predicted answer**
- D. Indonesia

Why this answer? India overtook China as the world's most populous country in 2023.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world facts recur in every Lecturer sitting.

Question 97 General Knowledge

Which is the longest mountain range in the world?

- A. Himalayas
- B. Andes ✓ **Predicted answer**
- C. Rockies
- D. Alps

Why this answer? The Andes in South America form the longest mountain range.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: world geography facts recur in every Lecturer sitting.

Question 98 General Knowledge

Which metal is liquid at room temperature?

- A. Mercury ✓ **Predicted answer**
- B. Iron
- C. Copper
- D. Aluminium

swer? Mercury is the only metal that is liquid at room temperature.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: everyday science facts recur in every Lecturer sitting.

Question 99 General Knowledge

Which planet is known as the Red Planet?

- A. Venus
- B. Mars ✓ **Predicted answer**
- C. Jupiter
- D. Saturn

Why this answer? Mars is called the Red Planet because of its iron-oxide surface.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: science facts recur in every Lecturer sitting.

Question 100 General Knowledge

Which gas do plants absorb from the air?

- A. Oxygen
- B. Nitrogen
- C. Hydrogen
- D. Carbon dioxide ✓ **Predicted answer**

Why this answer? Plants absorb carbon dioxide for photosynthesis.

Why predicted [C] — Syllabus anchor: syllabus anchor — General Knowledge: science facts recur in every Lecturer sitting.

This is a Ranjha Solutions prediction product, built from Ranjha Solutions' verified past-paper archive. It is **not** a PPSC paper, and no question here has appeared in any exam. Do not quote any item as "asked in the 2026 exam".

s are **free** — no charge is taken from any student. **No pass guarantee:** nobody can guarantee which questions will appear.