PPSC Lecturer Chemistry | BS-17 — GUESS PAPER 2026 (Set 1 of 2) · Predicted Questions

Body: Punjab Public Service Commission (PPSC) · **Post:** Lecturer Chemistry | BS-17 · **Set:** 1 of 2 · **Questions:** 100 predicted — 0 repeated questions, 33 verified facts predicted to recur, 67 syllabus anchors · **Price:** Free

GUESS PAPER — predicted questions for the upcoming exam, NOT a past paper. Every question below is a prediction: it is here because the same question, the same fact, or the same topic appeared in verified PPSC Lecturer Chemistry | BS-17 past papers. **No pass guarantee.** A guess paper raises the odds; it does not promise a result. Answers are the best-supported answers, not exam keys.

[C]-HEAVY PAPER — thin verified pool. 67 of the 100 questions below are syllabus anchors on this post's recurring topics, because the verified archive for this post holds only 66 items. Those questions are flagged [C] on every one of them: they are predictions about the syllabus, not verified facts. Study the topic, not just the answer.

> **How to use this paper:** study the basis under each question, not just the answer — the real paper will ask the same topics in different words. Revise the underlying topic, not only the four options. Do not quote any item as "asked in the 2026 exam" — nothing here has been asked yet.

Question 1 Chemistry

How many moles are there in 3 g of carbon (molar mass 12 g/mol)?

- A. 0.5
- B. 0.25 ✓ **Predicted answer**
- C. 0.75

Why this answer? Moles = mass / molar mass = 3 / 12 = 0.25 mol.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q1 (verified item — correct answer: 0.25).

Question 2 Chemistry

When diluting a concentrated acid, what is the safe order of mixing?

- A. Water is added in acid
- B. Either order is safe
- C. None of these
- D. Acid is added in water ✓ **Predicted answer**

Why this answer? Acid is added slowly to water so the heat of dilution is absorbed safely; adding water to acid causes violent spattering.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q2 (verified item — correct answer: Acid is added in water).

Question 3 Chemistry

Which of the following forms the strongest hydrogen bond?

- A. HF ✓ **Predicted answer**
- B. Water
- C. NH3
- D. None of these

Why this answer? The H-F bond is the most polar, so HF hydrogen bonding is the strongest.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q3 (verified item — correct answer: HF).

4 Chemistry

Boiling and melting points of a substance are related to the ___ of its particles.

- A. Size
- B. Distance between atoms
- C. None of these
- D. Kinetic energy ✓ **Predicted answer**

Why this answer? Boiling and melting points are related to the kinetic energy of the particles.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q4 (verified item — correct answer: Kinetic energy).

Question 5 Chemistry

Generating electricity by heating crystals is called:

- A. Piezo electricity
- B. Ferro electricity
- C. None of these
- D. Pyro electricity ✓ **Predicted answer**

Why this answer? Pyroelectricity is the generation of electricity by heating crystals.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q5 (verified item — correct answer: Pyro electricity).

Question 6 Chemistry

How many ions make up a unit cell of BaCl2?

- A. 4 Ba ions and 8 Cl ions ✓ **Predicted answer**
- B. 4 Ba ions and 4 Cl ions
- C. 8 Ba ions and 4 Cl ions

Why this answer? The BaCl2 (fluorite-type) unit cell contains 4 Ba ions and 8 Cl ions.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q6 (verified item — correct answer: 4 Ba ions and 8 Cl ions).

Question 7 Chemistry

Distances between atoms in molecular crystals are studied by which technique?

- A. Neutron spectroscopy
- B. Infrared spectroscopy
- C. X-ray crystallography ✓ **Predicted answer**
- D. Mass spectrometry

Why this answer? X-ray crystallography determines distances between atoms in crystals.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q7 (verified item — correct answer: X-ray crystallography).

Question 8 Chemistry

How does the energy of the 3d subshell compare with that of the 3p subshell?

- A. Smaller
- B. Greater ✓ **Predicted answer**
- C. Equal
- D. None of these

Why this answer? The 3d subshell has greater energy than the 3p subshell.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q9 (verified item — correct answer: Greater).

9 Chemistry

Which of the following is an example of electromagnetic radiation?

- A. Alpha
- B. Gamma ✓ **Predicted answer**
- C. Beta
- D. None of these

Why this answer? Gamma rays are electromagnetic radiation; alpha and beta are particles.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q10 (verified item — correct answer: Gamma).

Question 10 Chemistry

Which substance is a conductor and has covalent bonding?

- A. Nitrogen
- B. Graphite ✓ **Predicted answer**
- C. Oxygen
- D. None of these

Why this answer? Graphite conducts electricity and is covalently bonded.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q11 (verified item — correct answer: Graphite).

Question 11 Chemistry

Which molecule has a pyramidal structure?

- A. PCl3 ✓ **Predicted answer**
- B. H2O
- C. SO3

Why this answer? PCl3 has a pyramidal structure.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q12 (verified item — correct answer: PCl3).

Question 12 Chemistry

Which pair of species are iso-structural?

- A. NH3 and H2O
- B. CO2 and SO2
- C. XeF2 and IF2- ✓ **Predicted answer**
- D. None of these

Why this answer? XeF2 and IF2- are iso-structural.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q13 (verified item — correct answer: XeF2 and IF2-).

Question 13 Chemistry

Which species contains a coordinate covalent (dative) bond?

- A. NH4+ ✓ **Predicted answer**
- B. NH3
- C. H2O
- D. None of these

Why this answer? NH4+ contains a coordinate covalent bond.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q14 (verified item — correct answer: NH4+).

14 Chemistry

At equilibrium for A + B forming C + D, which statement is true?

- A. Reactants and products are present in equal amounts
- B. Reactants and products are present, generally at different concentrations ✓ **Predicted answer**
- C. Only products remain
- D. None of these

Why this answer? At equilibrium both reactants and products are present, generally at different concentrations.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q16 (verified item — correct answer: A, B and C, D are present at different concentration).

Question 15 Chemistry

If Kc = 2.5 for a reaction, the equilibrium lies:

- A. To the left
- B. To the right ✓ **Predicted answer**
- C. At equal concentrations
- D. None of these

Why this answer? Kc greater than 1 means the equilibrium lies to the right.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q17 (verified item — correct answer: Reaction lies at right).

Question 16 Chemistry

What does a high value of Kc indicate?

- A. The equilibrium lies to the left

reaction stops

- C. None of these
- D. The equilibrium lies far to the right ✓ **Predicted answer**

Why this answer? A high Kc means the reaction lies to the right and goes nearly to completion.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q18 (verified item — correct answer: Reaction is at right and going to be complete).

Question 17 Chemistry

What happens at chemical equilibrium?

- A. The reaction ceases
- B. The reaction reverses
- C. No net change occurs ✓ **Predicted answer**
- D. None of these

Why this answer? At equilibrium the forward and reverse rates are equal, so no net change occurs.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q19 (verified item — correct answer: No net change occurs).

Question 18 Chemistry

One milli-equivalent is equal to:

- A. 100 equivalents
- B. 10000 equivalents
- C. None of these
- D. 1/1000 equivalent ✓ **Predicted answer**

Why this answer? A milli-equivalent is 1/1000 of an equivalent.

ed [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q20 (verified item — correct answer: 1/1000 equivalent).

Question 19 Chemistry

One part per billion (ppb) is equal to:

- A. 1 nanogram per litre
- B. 1 milligram per litre
- C. 1 microgram per kilogram ✓ **Predicted answer**
- D. None of these

Why this answer? 1 ppb = 1 microgram per kilogram.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q21 (verified item — correct answer: microgram per kilogram).

Question 20 Chemistry

8 mL of a 0.25 M solution is diluted to 0.5 L. What is the new concentration?

- A. 0.25 M
- B. 8 M
- C. None of these
- D. 0.004 M ✓ **Predicted answer**

Why this answer? M2 = M1V1/V2 = 0.25 x 0.008 / 0.5 = 0.004 M.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q22 (verified item — correct answer: 0.004).

Question 21 Chemistry

A solution has 10 kg of solvent at a density of 5 kg per cubic decimetre and a concentration of 4 M. How much sodium (in grams) does it contain?

- B. 184 g ✓ **Predicted answer**
- C. 92 g
- D. None of these

Why this answer? Volume = 10/5 = 2 dm cubed; moles = 4 x 2 = 8; mass of Na = 8 x 23 = 184 g.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q23 (verified item — correct answer: 184).

Question 22 Chemistry

In expanded polystyrene foam, the dispersion medium is:

- A. Liquid
- B. Gas
- C. Solid ✓ **Predicted answer**
- D. None of these

Why this answer? In expanded polystyrene, gas is dispersed in a solid medium.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q24 (verified item — correct answer: Solid).

Question 23 Chemistry

What kind of mixture is fog?

- A. Liquid in gas ✓ **Predicted answer**
- B. Liquid in liquid
- C. Liquid in solid
- D. None of these

Why this answer? Fog is liquid droplets dispersed in a gas.

ed [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q25 (verified item — correct answer: Liquid in gas).

Question 24 Chemistry

What is the standard potential of the standard hydrogen electrode at pH = 10?

- A. 0 V
- B. -0.59 V ✓ **Predicted answer**
- C. +0.59 V
- D. None of these

Why this answer? E = -0.059 x pH = -0.59 V at pH 10.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q26 (verified item — correct answer: -0.59).

Question 25 Chemistry

A sample has a half-life of 30 minutes and starts with 600 atoms. How many atoms remain after 90 minutes?

- A. 150
- B. 75 ✓ **Predicted answer**
- C. 300
- D. None of these

Why this answer? 90 minutes is 3 half-lives: 600 / 2 cubed = 75 atoms.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q27 (verified item — correct answer: 75).

Question 26 Chemistry

Which of the following reactions is second order?

- B. Both reactions ✓ **Predicted answer**
- C. Only H2 + I2 forming 2HI
- D. Neither reaction

Why this answer? Both the ester hydrolysis with NaOH and H2 + I2 are second-order reactions.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q28 (verified item — correct answer: Both).

Question 27 Chemistry

Which of the following has a non-zero standard heat of formation?

- A. Fe (s)
- B. None of these ✓ **Predicted answer**
- C. N2 (g)
- D. H2 (g)

Why this answer? Elements in their standard states have zero standard heat of formation, so none of these has a non-zero value.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q29 (verified item — correct answer: None of these).

Question 28 Chemistry

Which equation represents the Freundlich adsorption isotherm?

- A. x/m = kP^(1/n), n > 2
- B. x/m = kP^(1/n), n > 3
- C. None of these
- D. x/m = kP^(1/n), n > 1 ✓ **Predicted answer**

swer? The Freundlich isotherm is x/m = kP^(1/n) with n greater than 1.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q30 (verified item — correct answer: x/m = kP1/n, n > 1).

Question 29 Chemistry

How is the quantum efficiency of a photochemical reaction determined?

- A. The number of photons emitted
- B. The number of molecules reacting and the number of photons absorbed ✓ **Predicted answer**
- C. The number of molecules colliding only
- D. None of these

Why this answer? Quantum efficiency is the number of molecules reacting per photon absorbed.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q31 (verified item — correct answer: Number of molecules colliding and number of photons absorbed).

Question 30 Chemistry

Which range of wavelengths can carry out a photochemical reaction?

- A. 500 to 1000 angstrom
- B. 1000 to 1500 angstrom
- C. None of these
- D. 2000 to 12000 angstrom ✓ **Predicted answer**

Why this answer? Photochemical reactions are carried out by radiation of about 2000 to 12000 angstrom.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-

Question 31 Chemistry

The electronic configuration 1s2 2s2 2p6 3s2 3p5 belongs to an element that is:

- A. Abundant
- B. Solid
- C. None of these
- D. Reactive ✓ **Predicted answer**

Why this answer? The configuration 1s2 2s2 2p6 3s2 3p5 is chlorine, a reactive element.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q34 (verified item — correct answer: Reactive).

Question 32 Chemistry

Which hydroxide is amphoteric?

- A. Pb(OH)2 ✓ **Predicted answer**
- B. NaOH
- C. KOH
- D. None of these

Why this answer? Lead(II) hydroxide is amphoteric.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q35 (verified item — correct answer: Pb(OH)2).

Question 33 Chemistry

Zinc does not show variable valency because of its:

- A. Inert pair effect
- B. 4s2 subshell

- D. None of these

Why this answer? Zinc has a completely filled d subshell, so it does not show variable valency.

Why predicted [B] — Verified fact, predicted to recur: ppsc-lecturer-chemistry-male-2022 Q37 (verified item — correct answer: Completely filled d subshell).

Question 34 Chemistry

How many moles are there in 6 g of a substance of molar mass 2 g/mol?

- A. 4
- B. 3 ✓ **Predicted answer**
- C. 2
- D. 6

Why this answer? Moles = mass / molar mass = 6 / 2 = 3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

Question 35 Chemistry

How many moles are there in 10 g of a substance of molar mass 5 g/mol?

- A. 3
- B. 1
- C. 2 ✓ **Predicted answer**
- D. 4

Why this answer? Moles = mass / molar mass = 10 / 5 = 2.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

36 Chemistry

How many moles are there in 12 g of a substance of molar mass 3 g/mol?

- A. 5
- B. 3
- C. 8
- D. 4 ✓ **Predicted answer**

Why this answer? Moles = mass / molar mass = 12 / 3 = 4.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

Question 37 Chemistry

How many moles are there in 20 g of a substance of molar mass 4 g/mol?

- A. 5 ✓ **Predicted answer**
- B. 6
- C. 4
- D. 10

Why this answer? Moles = mass / molar mass = 20 / 4 = 5.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: the mole concept recurs in every Lecturer Chemistry sitting.

Question 38 Chemistry

What is the molarity of a solution containing 2 moles of solute in 1 litres?

- A. 3 M
- B. 1 M
- C. 2 M ✓ **Predicted answer**
- D. 4 M

swer? Molarity = moles / volume = 2 / 1 = 2 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 39 Chemistry

What is the molarity of a solution containing 6 moles of solute in 2 litres?

- A. 3 M ✓ **Predicted answer**
- B. 4 M
- C. 2 M
- D. 5 M

Why this answer? Molarity = moles / volume = 6 / 2 = 3 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 40 Chemistry

What is the molarity of a solution containing 3 moles of solute in 1 litres?

- A. 4 M
- B. 3 M ✓ **Predicted answer**
- C. 2 M
- D. 5 M

Why this answer? Molarity = moles / volume = 3 / 1 = 3 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 41 Chemistry

What is the molarity of a solution containing 8 moles of solute in 4 litres?

- B. 2 M ✓ **Predicted answer**
- C. 1 M
- D. 4 M

Why this answer? Molarity = moles / volume = 8 / 4 = 2 M.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molarity and solutions recurs in every Lecturer Chemistry sitting.

Question 42 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-1 M?

- A. 1 ✓ **Predicted answer**
- B. 2
- C. 0
- D. 3

Why this answer? pH = -log[H+] = -log(10^-1) = 1.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 43 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-2 M?

- A. 3
- B. 1
- C. 2 ✓ **Predicted answer**
- D. 4

Why this answer? pH = -log[H+] = -log(10^-2) = 2.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs

urer Chemistry sitting.

Question 44 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-3 M?

- A. 3 ✓ **Predicted answer**
- B. 4
- C. 2
- D. 5

Why this answer? pH = -log[H+] = -log(10^-3) = 3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 45 Chemistry

What is the pH of a solution with hydrogen ion concentration 10^-4 M?

- A. 5
- B. 3
- C. 6
- D. 4 ✓ **Predicted answer**

Why this answer? pH = -log[H+] = -log(10^-4) = 4.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 46 Chemistry

A gas at 2 atm occupies 10 L. At constant temperature its volume becomes 5 L. What is the new pressure?

- A. 5 atm

m ✓ **Predicted answer**

- C. 3 atm
- D. 6 atm

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 2 x 10 / 5 = 4 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 47 Chemistry

A gas at 4 atm occupies 6 L. At constant temperature its volume becomes 3 L. What is the new pressure?

- A. 9 atm
- B. 8 atm ✓ **Predicted answer**
- C. 7 atm
- D. 10 atm

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 4 x 6 / 3 = 8 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 48 Chemistry

A gas at 3 atm occupies 8 L. At constant temperature its volume becomes 4 L. What is the new pressure?

- A. 7 atm
- B. 5 atm
- C. 6 atm ✓ **Predicted answer**
- D. 8 atm

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 3 x 8 / 4 = 6 atm.

ed [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 49 Chemistry

A gas at 6 atm occupies 12 L. At constant temperature its volume becomes 4 L. What is the new pressure?

- A. 19 atm
- B. 17 atm
- C. 18 atm ✓ **Predicted answer**
- D. 20 atm

Why this answer? By Boyle's law, P1V1 = P2V2, so P2 = 6 x 12 / 4 = 18 atm.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gas laws recurs in every Lecturer Chemistry sitting.

Question 50 Chemistry

What is the molar mass of H2SO4?

- A. 96 g/mol
- B. 98 g/mol ✓ **Predicted answer**
- C. 100 g/mol
- D. 80 g/mol

Why this answer? The molar mass of H2SO4 is 98 g/mol.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molar mass recurs in every Lecturer Chemistry sitting.

Question 51 Chemistry

What is the molar mass of NaOH?

- B. 40 g/mol ✓ **Predicted answer**
- C. 41 g/mol
- D. 36 g/mol

Why this answer? The molar mass of NaOH is 40 g/mol.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molar mass recurs in every Lecturer Chemistry sitting.

Question 52 Chemistry

What is the molar mass of CaCO3?

- A. 98 g/mol
- B. 102 g/mol
- C. 100 g/mol ✓ **Predicted answer**
- D. 84 g/mol

Why this answer? The molar mass of CaCO3 is 100 g/mol.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: molar mass recurs in every Lecturer Chemistry sitting.

Question 53 Chemistry

Which element has the chemical symbol Na?

- A. Nitrogen
- B. Sodium ✓ **Predicted answer**
- C. Nickel
- D. Neon

Why this answer? Na is the symbol for sodium.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and

Question 54 Chemistry

Which element has the chemical symbol Fe?

- A. Fluorine
- B. Iron ✓ **Predicted answer**
- C. Francium
- D. Fermium

Why this answer? Fe is the symbol for iron.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

Question 55 Chemistry

Which element is the lightest?

- A. Hydrogen ✓ **Predicted answer**
- B. Helium
- C. Oxygen
- D. Nitrogen

Why this answer? Hydrogen is the lightest element.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic table recurs in every Lecturer Chemistry sitting.

Question 56 Chemistry

Which gas is used in balloons because it is lighter than air?

- A. Helium ✓ **Predicted answer**
- B. Oxygen

ogen

- D. Argon

Why this answer? Helium is lighter than air and is used in balloons.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gases recurs in every Lecturer Chemistry sitting.

Question 57 Chemistry

Which acid is present in vinegar?

- A. Acetic acid ✓ **Predicted answer**
- B. Citric acid
- C. Sulphuric acid
- D. Nitric acid

Why this answer? Vinegar contains acetic acid.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

Question 58 Chemistry

Which acid is present in the stomach?

- A. Sulphuric acid
- B. Nitric acid
- C. Acetic acid
- D. Hydrochloric acid ✓ **Predicted answer**

Why this answer? The stomach contains hydrochloric acid.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

59 Chemistry

What is the pH of a neutral solution at 25 degrees Celsius?

- A. 7 ✓ **Predicted answer**
- B. 0
- C. 14
- D. 1

Why this answer? A neutral solution has pH 7 at 25 degrees Celsius.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: pH and acids recurs in every Lecturer Chemistry sitting.

Question 60 Chemistry

Which of the following is a strong acid?

- A. Sulphuric acid ✓ **Predicted answer**
- B. Acetic acid
- C. Citric acid
- D. Carbonic acid

Why this answer? Sulphuric acid is a strong acid.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

Question 61 Chemistry

Which gas is released when an acid reacts with a metal?

- A. Oxygen
- B. Carbon dioxide
- C. Nitrogen
- D. Hydrogen ✓ **Predicted answer**

swer? Acids react with metals to release hydrogen gas.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: acids and bases recurs in every Lecturer Chemistry sitting.

Question 62 Chemistry

Which metal is the best conductor of electricity?

- A. Silver ✓ **Predicted answer**
- B. Copper
- C. Aluminium
- D. Iron

Why this answer? Silver is the best conductor of electricity.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: metals recurs in every Lecturer Chemistry sitting.

Question 63 Chemistry

Which metal has the lowest melting point?

- A. Sodium
- B. Potassium
- C. Mercury ✓ **Predicted answer**
- D. Gallium

Why this answer? Mercury has the lowest melting point of the metals and is liquid at room temperature.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: metals recurs in every Lecturer Chemistry sitting.

64 Chemistry

Which non-metal is a good conductor of electricity?

- A. Sulphur
- B. Phosphorus
- C. Diamond
- D. Graphite ✓ **Predicted answer**

Why this answer? Graphite conducts electricity.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: non-metals recurs in every Lecturer Chemistry sitting.

Question 65 Chemistry

What is the chemical formula of common salt?

- A. NaCl ✓ **Predicted answer**
- B. KCl
- C. CaCl2
- D. NaOH

Why this answer? Common salt is sodium chloride, NaCl.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 66 Chemistry

What is the chemical formula of baking soda?

- A. Na2CO3
- B. NaHCO3 ✓ **Predicted answer**
- C. NaOH
- D. CaCO3

Answer: Baking soda is sodium hydrogen carbonate, NaHCO3.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 67 Chemistry

What is the chemical formula of limewater?

- A. CaO
- B. CaCO3
- C. CaCl2
- D. Ca(OH)2 ✓ **Predicted answer**

Why this answer? Limewater is calcium hydroxide, Ca(OH)2.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 68 Chemistry

Which gas is called laughing gas?

- A. Nitric oxide
- B. Carbon monoxide
- C. Sulphur dioxide
- D. Nitrous oxide ✓ **Predicted answer**

Why this answer? Nitrous oxide (N2O) is laughing gas.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: gases recurs in every Lecturer Chemistry sitting.

Question 69 Chemistry

Which element has the highest electronegativity?

orine ✓ **Predicted answer**

- B. Oxygen
- C. Chlorine
- D. Nitrogen

Why this answer? Fluorine has the highest electronegativity.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic trends recurs in every Lecturer Chemistry sitting.

Question 70 Chemistry

Which group of the periodic table contains the noble gases?

- A. Group 17
- B. Group 1
- C. Group 2
- D. Group 18 ✓ **Predicted answer**

Why this answer? The noble gases are in group 18.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic table recurs in every Lecturer Chemistry sitting.

Question 71 Chemistry

Which of the following is a noble gas?

- A. Chlorine
- B. Argon ✓ **Predicted answer**
- C. Oxygen
- D. Nitrogen

Why this answer? Argon is a noble gas.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: periodic table

Question 72 Chemistry

What is the atomic number of carbon?

- A. 12
- B. 4
- C. 6 ✓ **Predicted answer**
- D. 8

Why this answer? Carbon has atomic number 6.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 73 Chemistry

How many electrons does a neutral oxygen atom have?

- A. 6
- B. 16
- C. 10
- D. 8 ✓ **Predicted answer**

Why this answer? Oxygen has atomic number 8, so a neutral atom has 8 electrons.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 74 Chemistry

Which particle has a negative charge?

- A. Proton
- B. Neutron

- D. Electron ✓ **Predicted answer**

Why this answer? The electron carries a negative charge.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 75 Chemistry

Which particle has no charge?

- A. Electron
- B. Proton
- C. Neutron ✓ **Predicted answer**
- D. Positron

Why this answer? The neutron is uncharged.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: atomic structure recurs in every Lecturer Chemistry sitting.

Question 76 Chemistry

What is the chemical symbol for potassium?

- A. P
- B. K ✓ **Predicted answer**
- C. Pt
- D. Po

Why this answer? K (from kalium) is the symbol for potassium.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

77 Chemistry

What is the chemical symbol for calcium?

- A. C
- B. Cl
- C. Ca ✓ **Predicted answer**
- D. Cd

Why this answer? Ca is the symbol for calcium.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: elements and symbols recurs in every Lecturer Chemistry sitting.

Question 78 Chemistry

Which element is used in the filament of electric bulbs?

- A. Tungsten ✓ **Predicted answer**
- B. Copper
- C. Iron
- D. Nickel

Why this answer? Tungsten is used in bulb filaments because of its high melting point.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: metals recurs in every Lecturer Chemistry sitting.

Question 79 Chemistry

What is the chemical formula of sulphuric acid?

- A. HCl
- B. HNO3
- C. H2SO4 ✓ **Predicted answer**
- D. H2CO3

swer? Sulphuric acid is H2SO4.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: common compounds recurs in every Lecturer Chemistry sitting.

Question 80 Chemistry

Which gas is used with nitrogen in the Haber process to make ammonia?

- A. Oxygen
- B. Carbon dioxide
- C. Hydrogen ✓ **Predicted answer**
- D. Helium

Why this answer? Ammonia is made from nitrogen and hydrogen by the Haber process.

Why predicted [C] — Syllabus anchor: syllabus anchor — Chemistry: industrial chemistry recurs in every Lecturer Chemistry sitting.

Question 81 Islamiat

The shortest Surah of the Holy Quran is:

- A. Surah Al-Kawthar ✓ **Predicted answer**
- B. Surah Al-Ikhlas
- C. Surah Al-Asr
- D. Surah Al-Fatiha

Why this answer? Surah Al-Kawthar (3 verses) is the shortest Surah.

Why predicted [C] — Syllabus anchor: syllabus anchor — Islamiat: the Quran recurs in every Lecturer sitting.

Question 82 Pakistan Studies

The Lahore Resolution was passed in:

0 ✓ **Predicted answer**

- B. 1930
- C. 1947
- D. 1935

Why this answer? The Lahore Resolution was passed on 23 March 1940.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: the Pakistan Movement recurs in every Lecturer sitting.

Question 83 Pakistan Studies

The first Constitution of Pakistan was adopted in:

- A. 1947
- B. 1962
- C. 1956 ✓ **Predicted answer**
- D. 1973

Why this answer? The first Constitution of Pakistan was adopted in 1956.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: constitutional history recurs in every Lecturer sitting.

Question 84 Pakistan Studies

The current Constitution of Pakistan dates from:

- A. 1956
- B. 1962
- C. 1985
- D. 1973 ✓ **Predicted answer**

Why this answer? The current Constitution of Pakistan was adopted in 1973.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: constitutional

s in every Lecturer sitting.

Question 85 Pakistan Studies

Who presented the Lahore Resolution?

- A. Muhammad Ali Jinnah
- B. Liaquat Ali Khan
- C. Allama Iqbal
- D. A. K. Fazlul Huq ✓ **Predicted answer**

Why this answer? A. K. Fazlul Huq presented the Lahore Resolution in 1940.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: the Pakistan Movement recurs in every Lecturer sitting.

Question 86 Pakistan Studies

The Durand Line is the border between Pakistan and:

- A. Iran
- B. India
- C. China
- D. Afghanistan ✓ **Predicted answer**

Why this answer? The Durand Line is the Pakistan-Afghanistan border.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: Pakistan geography recurs in every Lecturer sitting.

Question 87 Pakistan Studies

The Simla Agreement was signed in:

- A. 1965
- B. 1971

2 ✓ **Predicted answer**

- D. 1974

Why this answer? The Simla Agreement was signed in 1972.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: Pakistan history recurs in every Lecturer sitting.

Question 88 Pakistan Studies

Who was the first Prime Minister of Pakistan?

- A. Muhammad Ali Jinnah
- B. Khwaja Nazimuddin
- C. Liaquat Ali Khan ✓ **Predicted answer**
- D. Ayub Khan

Why this answer? Liaquat Ali Khan was the first Prime Minister of Pakistan.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: Pakistan history recurs in every Lecturer sitting.

Question 89 Pakistan Studies

The Objective Resolution was passed in:

- A. 1949 ✓ **Predicted answer**
- B. 1947
- C. 1956
- D. 1973

Why this answer? The Objective Resolution was passed in March 1949.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: constitutional history recurs in every Lecturer sitting.

90 Pakistan Studies

The largest dam in Pakistan is:

- A. Tarbela Dam ✓ **Predicted answer**
- B. Mangla Dam
- C. Warsak Dam
- D. Diamer-Bhasha Dam

Why this answer? Tarbela is the largest dam in Pakistan.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: Pakistan geography recurs in every Lecturer sitting.

Question 91 Pakistan Studies

Which city of Pakistan is known as the City of Saints?

- A. Lahore
- B. Peshawar
- C. Multan ✓ **Predicted answer**
- D. Quetta

Why this answer? Multan is known as the City of Saints.

Why predicted [C] — Syllabus anchor: syllabus anchor — Pakistan Studies: Pakistan history recurs in every Lecturer sitting.

Question 92 Everyday Science

The chemical formula of water is:

- A. CO2
- B. O2
- C. NaCl
- D. H2O ✓ **Predicted answer**

swer? Water is H2O.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: basic chemistry facts recur in every Lecturer sitting.

Question 93 Everyday Science

The speed of light in a vacuum is about:

- A. 150,000 km per second
- B. 30,000 km per second
- C. 300,000 km per second ✓ **Predicted answer**
- D. 3,000 km per second

Why this answer? Light travels at about 300,000 km per second in a vacuum.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: physics facts recur in every Lecturer sitting.

Question 94 Everyday Science

Which vitamin is found in citrus fruits?

- A. Vitamin A
- B. Vitamin D
- C. Vitamin B12
- D. Vitamin C ✓ **Predicted answer**

Why this answer? Citrus fruits are rich in vitamin C.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: nutrition facts recur in every Lecturer sitting.

Question 95 Everyday Science

The boiling point of water at sea level is:

- B. 80 degrees Celsius
- C. 120 degrees Celsius
- D. 100 degrees Celsius ✓ **Predicted answer**

Why this answer? Water boils at 100 degrees Celsius at sea level.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: basic physics facts recur in every Lecturer sitting.

Question 96 Everyday Science

Which blood cells carry oxygen?

- A. White blood cells
- B. Platelets
- C. Red blood cells ✓ **Predicted answer**
- D. Plasma cells

Why this answer? Red blood cells carry oxygen through haemoglobin.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: human biology facts recur in every Lecturer sitting.

Question 97 Everyday Science

The SI unit of force is:

- A. Joule
- B. Newton ✓ **Predicted answer**
- C. Watt
- D. Pascal

Why this answer? The newton is the SI unit of force.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: units recur in

Question 98 Everyday Science

Which gas is most abundant in the air?

- A. Oxygen
- B. Carbon dioxide
- C. Nitrogen ✓ **Predicted answer**
- D. Argon

Why this answer? Nitrogen makes up about 78 percent of the air.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: basic chemistry facts recur in every Lecturer sitting.

Question 99 Everyday Science

The powerhouse of the cell is the:

- A. Mitochondrion ✓ **Predicted answer**
- B. Nucleus
- C. Ribosome
- D. Golgi body

Why this answer? The mitochondrion is the powerhouse of the cell.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: cell biology facts recur in every Lecturer sitting.

Question 100 Everyday Science

How many bones are there in the adult human body?

- A. 206 ✓ **Predicted answer**
- B. 196

- D. 226

Why this answer? An adult human has 206 bones.

Why predicted [C] — Syllabus anchor: syllabus anchor — Everyday Science: human biology facts recur in every Lecturer sitting.

This is a Ranjha Solutions prediction product, built from Ranjha Solutions' verified past-paper archive. It is **not** a PPSC paper, and no question here has appeared in any exam. Do not quote any item as "asked in the 2026 exam".

Guess papers are **free** — no charge is taken from any student. **No pass guarantee:** nobody can guarantee which questions will appear.